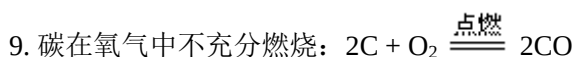
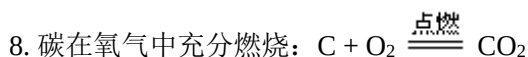
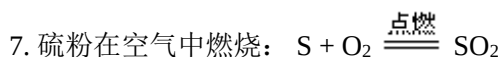
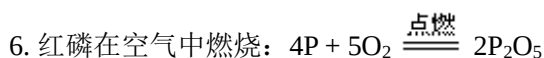
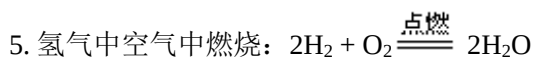
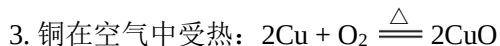
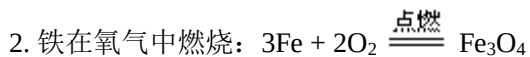
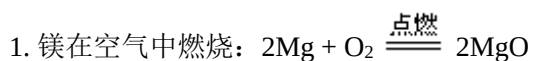


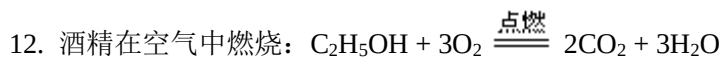
初中化学方程式分类汇总

一. 物质与氧气的反应:

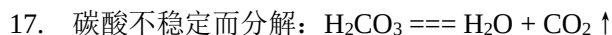
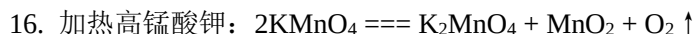
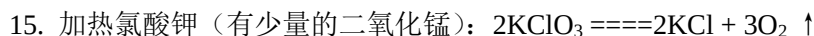
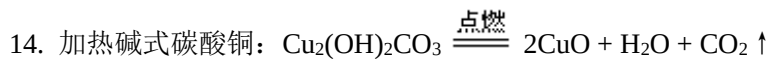
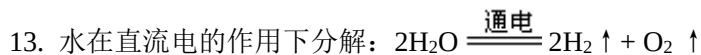
(1) 单质与氧气的反应:



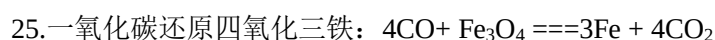
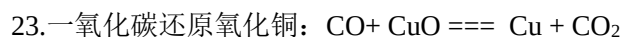
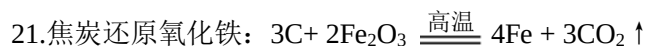
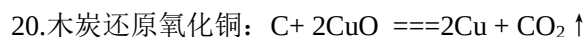
(2) 化合物与氧气的反应:



二. 几个分解反应:

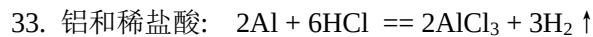
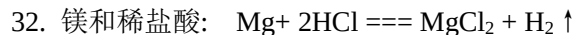
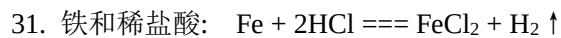
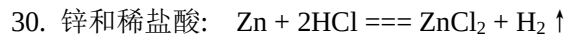
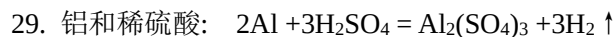
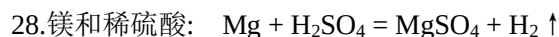
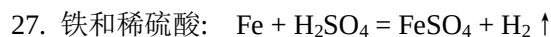
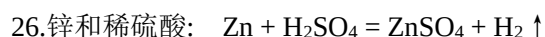


三. 几个氧化还原反应:

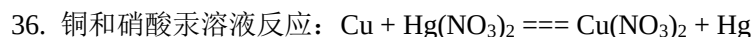
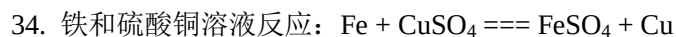


四. 单质、氧化物、酸、碱、盐的相互关系

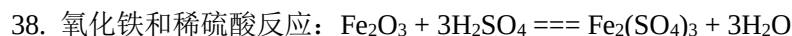
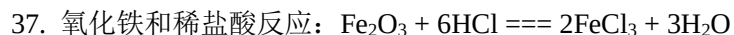
(1) 金属单质 + 酸 ----- 盐 + 氢气 (置换反应)



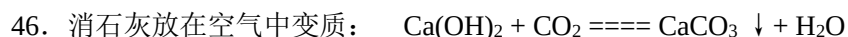
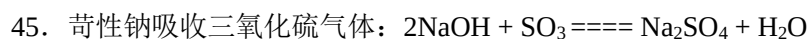
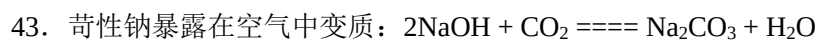
(2) 金属单质 + 盐 (溶液) ----- 另一种金属 + 另一种盐 (置换反应)



(3) 碱性氧化物 + 酸 ----- 盐 + 水 (复分解反应)



(4) 酸性氧化物 + 碱 ----- 盐 + 水 (非复分解反应)



47. 消石灰吸收二氧化硫: $\text{Ca}(\text{OH})_2 + \text{SO}_2 \rightleftharpoons \text{CaSO}_3 \downarrow + \text{H}_2\text{O}$

(5) 酸 + 碱 ----- 盐 + 水 (复分解反应)

48. 盐酸和烧碱起反应: $\text{HCl} + \text{NaOH} \rightleftharpoons \text{NaCl} + \text{H}_2\text{O}$
49. 盐酸和氢氧化钾反应: $\text{HCl} + \text{KOH} \rightleftharpoons \text{KCl} + \text{H}_2\text{O}$
50. 盐酸和氢氧化铜反应: $2\text{HCl} + \text{Cu}(\text{OH})_2 \rightleftharpoons \text{CuCl}_2 + 2\text{H}_2\text{O}$
51. 盐酸和氢氧化钙反应: $2\text{HCl} + \text{Ca}(\text{OH})_2 \rightleftharpoons \text{CaCl}_2 + 2\text{H}_2\text{O}$
52. 盐酸和氢氧化铁反应: $3\text{HCl} + \text{Fe}(\text{OH})_3 \rightleftharpoons \text{FeCl}_3 + 3\text{H}_2\text{O}$
53. 氢氧化铝药物治疗胃酸过多: $3\text{HCl} + \text{Al}(\text{OH})_3 \rightleftharpoons \text{AlCl}_3 + 3\text{H}_2\text{O}$
54. 硫酸和烧碱反应: $\text{H}_2\text{SO}_4 + 2\text{NaOH} \rightleftharpoons \text{Na}_2\text{SO}_4 + 2\text{H}_2\text{O}$
55. 硫酸和氢氧化钾反应: $\text{H}_2\text{SO}_4 + 2\text{KOH} \rightleftharpoons \text{K}_2\text{SO}_4 + 2\text{H}_2\text{O}$
56. 硫酸和氢氧化铜反应: $\text{H}_2\text{SO}_4 + \text{Cu}(\text{OH})_2 \rightleftharpoons \text{CuSO}_4 + 2\text{H}_2\text{O}$
57. 硫酸和氢氧化铁反应: $3\text{H}_2\text{SO}_4 + 2\text{Fe}(\text{OH})_3 \rightleftharpoons \text{Fe}_2(\text{SO}_4)_3 + 6\text{H}_2\text{O}$
58. 硝酸和烧碱反应: $\text{HNO}_3 + \text{NaOH} \rightleftharpoons \text{NaNO}_3 + \text{H}_2\text{O}$

(6) 酸 + 盐 ----- 另一种酸 + 另一种盐 (复分解反应)

59. 大理石与稀盐酸反应: $\text{CaCO}_3 + 2\text{HCl} \rightleftharpoons \text{CaCl}_2 + \text{H}_2\text{O} + \text{CO}_2 \uparrow$
60. 碳酸钠与稀盐酸反应: $\text{Na}_2\text{CO}_3 + 2\text{HCl} \rightleftharpoons 2\text{NaCl} + \text{H}_2\text{O} + \text{CO}_2 \uparrow$
61. 碳酸镁与稀盐酸反应: $\text{MgCO}_3 + 2\text{HCl} \rightleftharpoons \text{MgCl}_2 + \text{H}_2\text{O} + \text{CO}_2 \uparrow$
62. 盐酸和硝酸银溶液反应: $\text{HCl} + \text{AgNO}_3 \rightleftharpoons \text{AgCl} \downarrow + \text{HNO}_3$
63. 硫酸和碳酸钠反应: $\text{Na}_2\text{CO}_3 + \text{H}_2\text{SO}_4 \rightleftharpoons \text{Na}_2\text{SO}_4 + \text{H}_2\text{O} + \text{CO}_2 \uparrow$
64. 硫酸和氯化钡溶液反应: $\text{H}_2\text{SO}_4 + \text{BaCl}_2 \rightleftharpoons \text{BaSO}_4 \downarrow + 2\text{HCl}$

(7) 碱 + 盐 ----- 另一种碱 + 另一种盐 (复分解反应, 且两种反应物均可溶)

65. 氢氧化钠与硫酸铜: $2\text{NaOH} + \text{CuSO}_4 \rightleftharpoons \text{Cu}(\text{OH})_2 \downarrow + \text{Na}_2\text{SO}_4$
66. 氢氧化钠与氯化铁: $3\text{NaOH} + \text{FeCl}_3 \rightleftharpoons \text{Fe}(\text{OH})_3 \downarrow + 3\text{NaCl}$
67. 氢氧化钠与氯化镁: $2\text{NaOH} + \text{MgCl}_2 \rightleftharpoons \text{Mg}(\text{OH})_2 \downarrow + 2\text{NaCl}$
68. 氢氧化钠与氯化铜: $2\text{NaOH} + \text{CuCl}_2 \rightleftharpoons \text{Cu}(\text{OH})_2 \downarrow + 2\text{NaCl}$
69. 氢氧化钙与碳酸钠: $\text{Ca}(\text{OH})_2 + \text{Na}_2\text{CO}_3 \rightleftharpoons \text{CaCO}_3 \downarrow + 2\text{NaOH}$

(8) 盐 + 盐 ----- 新盐 1 + 新盐 2 (复分解反应, 且两种反应物均可溶)

70. 氯化钠溶液和硝酸银溶液: $\text{NaCl} + \text{AgNO}_3 \rightleftharpoons \text{AgCl} \downarrow + \text{NaNO}_3$
71. 硫酸钠和氯化钡: $\text{Na}_2\text{SO}_4 + \text{BaCl}_2 \rightleftharpoons \text{BaSO}_4 \downarrow + 2\text{NaCl}$

五. 其它反应:

72. 二氧化碳溶于水: $\text{CO}_2 + \text{H}_2\text{O} \rightleftharpoons \text{H}_2\text{CO}_3$
73. 生石灰溶于水: $\text{CaO} + \text{H}_2\text{O} \rightleftharpoons \text{Ca}(\text{OH})_2$
74. 氧化钠溶于水: $\text{Na}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons 2\text{NaOH}$
75. 三氧化硫溶于水: $\text{SO}_3 + \text{H}_2\text{O} \rightleftharpoons \text{H}_2\text{SO}_4$
76. 硫酸铜晶体受热分解: $\text{CuSO}_4 \cdot 5\text{H}_2\text{O} \rightleftharpoons \text{CuSO}_4 + 5\text{H}_2\text{O}$
77. 无水硫酸铜作干燥剂: $\text{CuSO}_4 + 5\text{H}_2\text{O} \rightleftharpoons \text{CuSO}_4 \cdot 5\text{H}_2\text{O}$